

# University of Pretoria Yearbook 2016

## General chemistry 117 (CMY 117)

**Qualification** Undergraduate

**Faculty** Faculty of Natural and Agricultural Sciences

**Module credits** 16.00

**Programmes** BEd Senior Phase and Further Education and Training Teaching

BSc(Computer Science) Computer Science

BDietetics Dietetics

BSc Biochemistry

BSc Biological Sciences

BSc Biotechnology

BSc Chemistry

BSc Ecology

BSc Entomology

BSc Environmental and Engineering Geology

BSc Environmental Sciences

BSc Food Management (4 years)

BSc Food Science

BSc Genetics

BSc Geography

BSc Geology

BSc Human Genetics

BSc Human Physiology

BSc Human Physiology, Genetics and Psychology

BSc Medical Sciences

BSc Meteorology

BSc Microbiology

BSc Nutrition

BSc Physics

BSc Plant Science



BSc Zoology

BScAgric Agricultural Economics: Agribusiness Management

BScAgric Animal Science

BScAgric Animal Science: Pasture Science

BScAgric Food Science and Technology

BScAgric Option: Applied Plant and Soil Sciences

BScAgric Plant Pathology

BVeterinary Science Veterinary Science

**Service modules** Faculty of Engineering, Built Environment and Information Technology

Faculty of Education

Faculty of Health Sciences

Faculty of Veterinary Science

**Prerequisites** Final Grade 12 marks of at least 60% for Mathematics and 60% for Physical Sciences.

**Contact time** 1 practical per week, 4 lectures per week

**Language of tuition** Both Afr and Eng

**Academic organisation** Chemistry

**Period of presentation** Semester 1

### Module content

General introduction to inorganic, analytical and physical chemistry. Atomic structure and periodicity. Molecular structure and chemical bonding using the VSEOR model. Nomenclature of iorganic ions and compounds. Classification of reactions: precipitation, acid-base, redox reactions and gas-forming reactions. Mole concept and stoichiometric calculations concerning chemical formulas and chemical reactions. Principles of reactivity: energy and chemical reactions. Physical behaviour gases, liquids, solids and solutions and the role of intermolecular forces. Rate of reactions: Introduction to chemical kinetics.

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